

## NAMES OF SIMPLE IONS

Cations:				Anions:	
	-OUS		-IC		-IDE
$\text{Cr}^{2+}$	chromium (II) chromous	$\text{Cr}^{3+}$	chromium (III) chromic	$\text{N}^{3-}$	nitride
$\text{Mn}^{2+}$	manganese (II) manganous	$\text{Mn}^{3+}$	manganese (III) manganic	$\text{O}^{2-}$	oxide
$\text{Fe}^{2+}$	iron (II) ferrous	$\text{Fe}^{3+}$	iron (III) ferric	$\text{S}^{2-}$	sulfide
$\text{Co}^{2+}$	cobalt (II) cobaltous	$\text{Co}^{3+}$	cobalt (III) cobaltic	$\text{H}^-$	hydride
$\text{Ni}^{2+}$	nickel (II) nickelous	$\text{Ni}^{3+}$	nickel (III) nickelic	$\text{F}^-$	fluoride
$\text{Cu}^+$	copper (I) cuprous	$\text{Cu}^{2+}$	copper (II) cupric	$\text{Cl}^-$	chloride
$\text{Au}^+$	gold (I) aurous	$\text{Au}^{3+}$	gold (III) auric	$\text{Br}^-$	bromide
$\text{Hg}_2^{2+}$	mercury (I) mercurous	$\text{Hg}^{2+}$	mercury (II) mercuric	$\text{I}^-$	iodide
$\text{Sn}^{2+}$	tin (II) stannous	$\text{Sn}^{4+}$	tin (IV) stannic	For all other <u>cations</u> , the name of the ion is the same as the name of the element from which it derives.	
$\text{Pb}^{2+}$	lead (II) plumbous	$\text{Pb}^{4+}$	lead (IV) plumbic		
$\text{As}^{3+}$	arsenic (III)	$\text{As}^{5+}$	arsenic (V)		
$\text{Sb}^{3+}$	antimony (III)	$\text{Sb}^{5+}$	antimony (V)		
$\text{Bi}^{3+}$	bismuth (III)	$\text{Bi}^{5+}$	bismuth (V)		