

Electron Configuration
CHEMISTRY 110

Name _____
last first .

Fill in the blanks:

Isotope	Number of protons	Number of neutrons	Electron configuration Starting with the 1s subshell
Beryllium-9	4	5	$1s^2 2s^2$
Aluminum-26	13	13	$1s^2 2s^2 2p^6 3s^2 3p^1$
Sodium-23	11	12	$1s^2 2s^2 2p^6 3s^1$
Tin-101	50	51	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^2$
Rhodium-107	45	62	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^7$
K ⁺ (Mass number = 43)	19	24	$1s^2 2s^2 2p^6 3s^2 3p^6$
Cl ⁻ (Mass number = 40)	17	23	$1s^2 2s^2 2p^6 3s^2 3p^6$
P ³⁻ (Mass number = 31)	15	16	$1s^2 2s^2 2p^6 3s^2 3p^6$
Mg ²⁺ (Mass number = 26)	12	14	$1s^2 2s^2 2p^6$