

**Nomenclature**  
**CHEMISTRY 110**

Name \_\_\_\_\_  
last first

1. Write the chemical formulas for the compounds formed between the following ions:

- a. Potassium-----bromide \_\_\_\_\_
- b. Ferric ion----Sulfide \_\_\_\_\_
- c. Tin(IV)----Phosphide \_\_\_\_\_
- d. Manganese (III)-----Sulfite \_\_\_\_\_
- e. Chromic---Oxalate \_\_\_\_\_
- f. Zinc-----Bromite \_\_\_\_\_
- g. Aluminum----iodate \_\_\_\_\_
- h. Cobaltous-----Bicarbonate \_\_\_\_\_
- i. Mercury (I)-----Nitrate \_\_\_\_\_
- j. Barium-----mono hydrogen phosphate \_\_\_\_\_

2. Predict the formula for the following compounds:

- a. Strontium chloride \_\_\_\_\_
- b. Sulfurous Acid \_\_\_\_\_
- c. Diarsenic Pentasulfide \_\_\_\_\_
- d. Cupric Hypochlorite \_\_\_\_\_
- e. Nickel (II) Phosphate \_\_\_\_\_
- f. Zinc Nitride \_\_\_\_\_
- g. Silver Dichromate \_\_\_\_\_
- h. Potassium Permanganate \_\_\_\_\_
- i. Carbon tetrachloride \_\_\_\_\_
- j. Periodic Acid \_\_\_\_\_

3. Name the following compounds:

- a.  $P_4O_6$  \_\_\_\_\_
- b.  $HBrO_4(aq)$  \_\_\_\_\_
- c.  $HBr(aq)$  \_\_\_\_\_
- d.  $HgCrO_4$  \_\_\_\_\_
- e.  $Sn(SO_3)_2$  \_\_\_\_\_
- f.  $OF_2$  \_\_\_\_\_
- g.  $AlN$  \_\_\_\_\_
- h.  $K_2O_2$  \_\_\_\_\_
- i.  $H_2CO_3(aq)$  \_\_\_\_\_
- j.  $SiO_4$  \_\_\_\_\_